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have different masses and temperatures. If the objects are brought into thermal contact, which one will have the greater temperature change?

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14-1 Heat As Energy Transfer . We often speak of heat as though it were a material that flows from one object to

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another; it is not.
Rather, it is a

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Summary of Chapter
14 • Internal energy U
refers to the total
energy of all molecules
in an object. For an
ideal monatomic gas,

- Heat is the transfer of

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energy from one object to another due to a temperature difference. Heat can be measured in joules or in calories. • Specific heat of a substance is the energy

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Equivalent of Heat. The

amount of work equal

to one calorie (4.186 J

$= 1 \text{ cal}$) Heat. Energy

(J) transferred from one

object to another

because of a difference

in temperature (always

a higher to lower

energy transfer)

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