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CHAPTER 18: Chemical Equilibrium I. a. Write the chemical equilibrium expression for the following equations. Include the value of K . (a) $(g) \rightleftharpoons (g)$ $K=0.1$ (b) $(g) + O_2(g) \rightleftharpoons (g)$ $K = 1.8 \times 10^{-2}$ Does the reaction favor products or reactants? (Will there be mostly products or reactants when it reaches 2. a.

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Compare the rates of forward and reverse reactions when equilibrium has been reached. b.

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SECTION 2 SHORT

ANSWER Answer the following questions in the space provided. 1.

_____ Raising the temperature of any equilibrium system always (a) favors the forward reaction. (b) favors the reverse reaction. (c) favors the

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exothermic reaction.

(d) favors the
endothermic reaction.

2. Consider the
following equilibrium
equation:

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19 chemistry; Chemical
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Equilibrium. reversible
reaction. chemical
equilibrium. law of
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state of balance. can occur in both the forward and reverse directions. state of balance in which forward and reverse reactions take p....

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The balanced chemical equation for an equilibrium system helps establish the expression for the equilibrium constant. The data in

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Table 18-1 show that the validity of this expression is confirmed when the actual values of the equilibrium concentrations of reactants and products are determined experimentally. The values of K_{eq} are calculated from these concentrations.

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CHEM-1412, Chapter
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This chemistry video tutorial focuses on the reaction quotient Q and the equilibrium constant K .

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1. In a bottle of unopened cola, the CO₂ gas dissolved in the liquid is in equilibrium with the CO₂ gas above the liquid. The dissolved gas reacts with water molecules in the cola to form carbonic acid, which also dissociates into carbon dioxide and

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water. Which chemical equation(s) best describe this equilibrium system? a. $\text{CO}_2(\text{g}) \rightleftharpoons \text{CO}_2(\text{l})$ b.

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